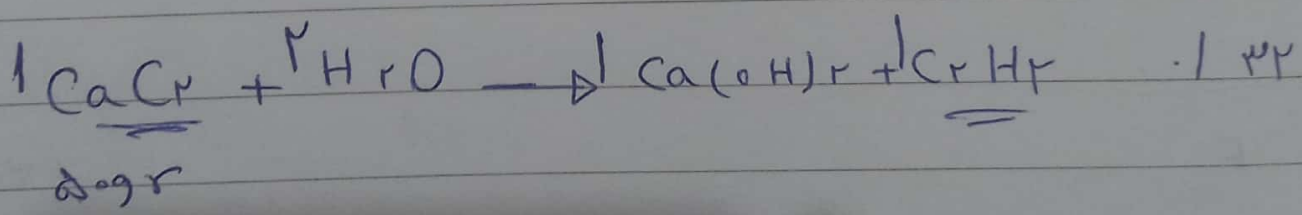


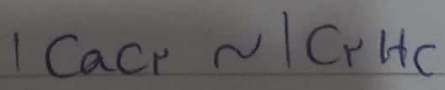
$\frac{1 \times 94.1 \times 10^3}{1 \times 342} = \frac{V}{V \times 22.4} \Rightarrow V = 14.4 \times 10^3$

$\frac{2 \times 101.6 \times 10^3}{2 \times 170} = \frac{V}{V \times 22.4} \Rightarrow V = 14.4 \times 10^3$



reactions P

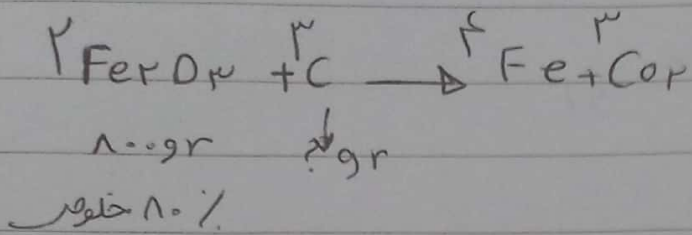
2P = 2P



$\frac{P \times 2P \times 50}{1 \times 100 \times 100} = \frac{4}{1 \times 342} \Rightarrow P = 14.4$
 $P = 8.1\%$

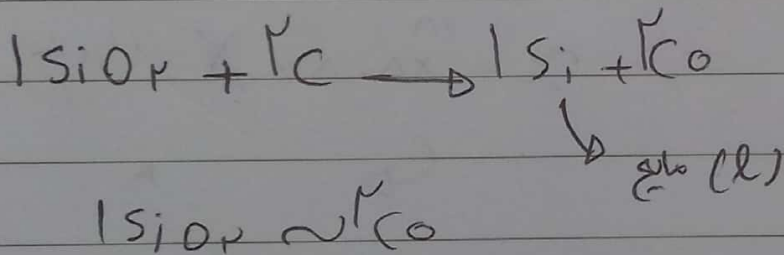
4.0% soluble

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$$\frac{1 \times 100 \times 100}{1 \times 100 \times 140} = \frac{9\%}{1 \times 14}$$

\rightarrow (100%) ✓

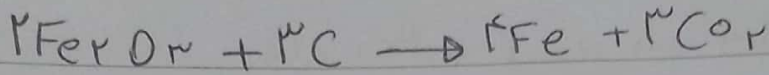


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$$\frac{48 \times 9,9}{100 \times 1 \times 41} = \frac{\lambda \text{ Lit}}{1 \times 22,1 \text{ E}} \rightarrow \lambda = 9,9$$

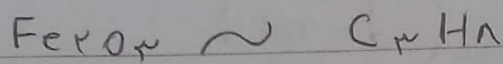
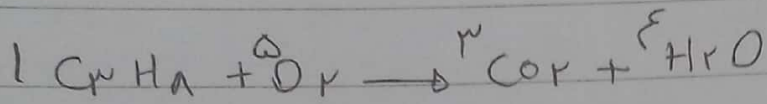
F

135

$\xi = 1.30$ 

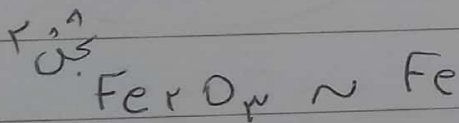
$$\downarrow$$

$$x \cdot 160$$

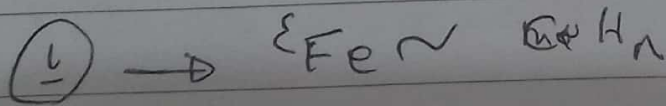


مساوية P

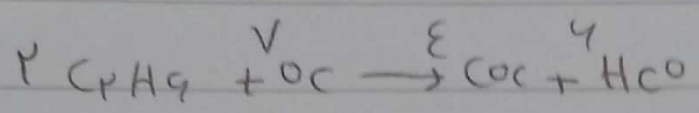
$$\frac{1 \cdot x \cdot 160 \cdot 10^{-3}}{1 \cdot x \cdot 14} = \frac{x \text{ mol}}{1} \rightarrow x = 1.00 \text{ mol}$$



$$\frac{P \cdot 160 \cdot 10^{-3}}{1 \cdot x \cdot 14} = \frac{149.15 \cdot 10^{-3}}{x \cdot 56} \rightarrow P = 1.0\%$$



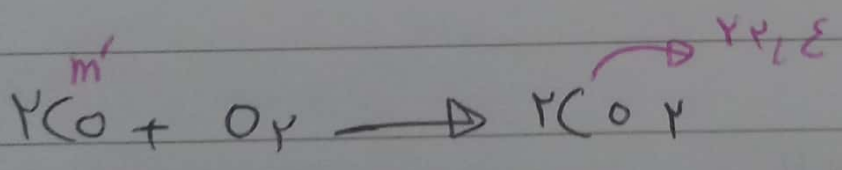
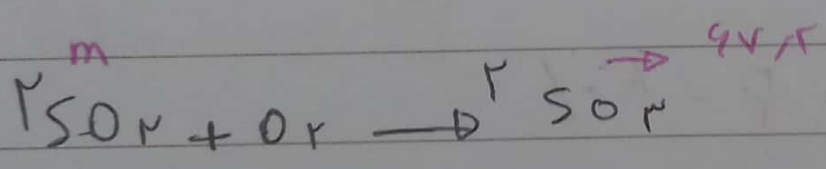
$$\frac{149.15 \cdot 10^{-3}}{x \cdot 56} = \frac{x \text{ mol}}{1} \rightarrow x = 1.00$$



$NaHCO_3 \sim CO_2$

$$\frac{n \times \Delta P / D}{1 \times 1 \times 1} = \frac{1 \times d_{CO_2}}{1 \times 1 \times 1} \rightarrow d = 1/1$$

$$\frac{m_{CO_2}}{1 \times 1} = \frac{n \times Lit}{1 \times 1 \times 1} \rightarrow V = V Lit$$



$SO_2 \sim SO_3$

$$\frac{x}{1 \times 4} = \frac{4V/5}{1 \times 2 \times 1} \rightarrow x = 195$$

CO

$CO \sim CO_2$

$$\frac{m}{1 \times 1} = \frac{2V/1}{1 \times 2 \times 1} \rightarrow m = 97.5$$

$$\Gamma_{SO_2} \sim b_{SO_2} = \frac{q}{\Delta H}$$

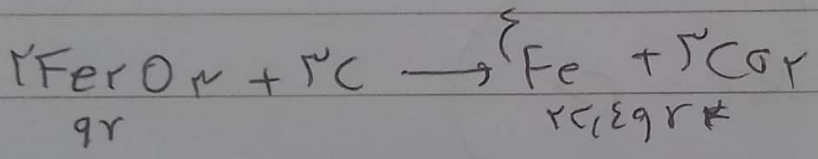
$$\frac{4 \times 10^{-4}}{1 \times 10^4} = \frac{q}{197} \rightarrow q = 7.98$$

$$\Gamma_{CO} \sim b_{CO}$$

$$\frac{1 \times 10^{-4}}{1 \times 10^4} = \frac{q}{0.99} \rightarrow q = 9.9 \times 10^{-9}$$

$\rightarrow \Delta V V$

\int^2 حل المسألة



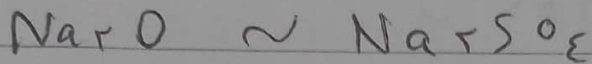
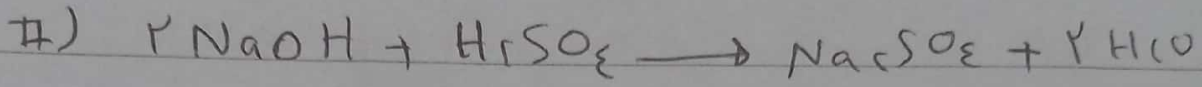
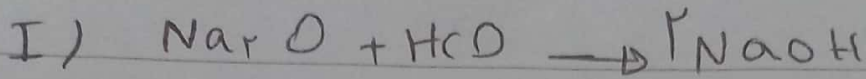
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$$\Gamma_{Fe_2O_3} \sim \Gamma_{Fe} \sim \Gamma_{CO_2}$$

$$\frac{x_{gr}}{1 \times 10^4} = \frac{1 \times 10^{-4}}{1 \times 10^4} = \frac{y_{gr}}{1 \times 10^4}$$

$$x = 10^{-4}$$

$$y = 10^{-4}$$



$$\frac{124}{1 \times 42} = \frac{296}{1 \times 142} \Rightarrow x = 12 \times 142$$

درصد جرمی Na_2SO_4 = $\frac{\text{جرم مولی } \text{Na}_2\text{SO}_4}{\text{جرم مولی محلول}} \times 100 = \frac{142 \times 12}{1200} \times 100 = 141.5\%$

عملیاتی \rightarrow $\text{جرم} = \text{حجم} \times \text{چگالی}$
 $1000 \times 1.2 = 1200$