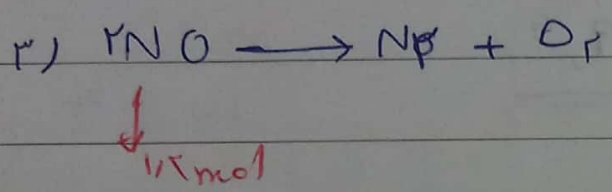
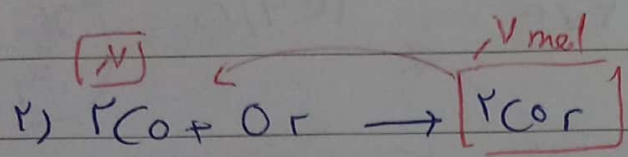
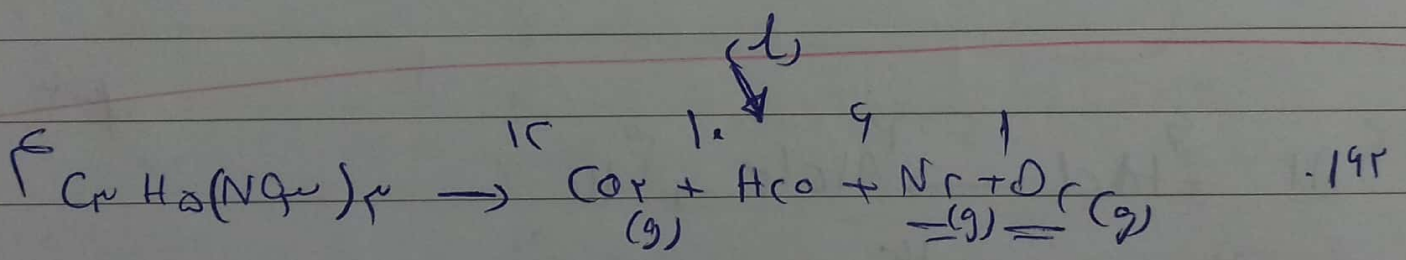


$\frac{2,9}{1,9} = 1,5 \text{ mol}$



$\frac{1,9}{2} \times 100 = 9,5\%$



دردهای VAC - گاز CO_2 جامد می شود! \leftarrow $\text{C}_2\text{H}_5(\text{NO}_2)_2$ \sim STP \leftarrow $\text{C}_2\text{H}_5(\text{NO}_2)_2$ \sim CO_2 \leftarrow $\text{C}_2\text{H}_5(\text{NO}_2)_2$ \sim CO_2

گازهای برف \sim $\frac{1 \text{ mol } \text{C}_2\text{H}_5(\text{NO}_2)_2}{19 \times 25,2}$

\leftarrow $104,8 \text{ Lit}$ \leftarrow CO_2

$\frac{1}{8} = \frac{\text{Lit } \text{CO}_2}{4 \text{ Vol } \text{Lit } \text{CO}_2} \rightarrow 9,75 \text{ Lit } \text{CO}_2$

$P_1 V_1 \rightarrow 1.9, \text{g} - 4V, \text{r} = 39, \text{r}$

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2} \rightarrow \frac{39, \text{r}}{270} = \frac{V_2}{129, \text{r}}$$

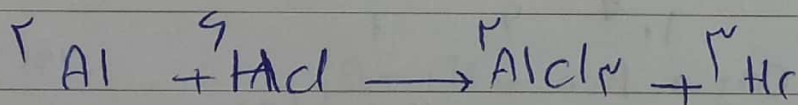
$V_2 = 19, \text{g}$

$- 129, \text{r} + 270 = 129, \text{r}$

$39, \text{r} - 19, \text{g} = 19, \text{g}$

$4V, \text{r} + 19, \text{g} = 19, \text{g}$

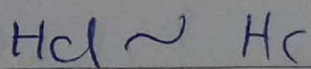
193



$$\frac{9}{2 \times 27} = \frac{V \times 10^{-3}}{\frac{V}{2} \times 2} \rightarrow V \times 10^{-3} = 2, 199 \text{r}$$

علاقة مولات $\rightarrow \frac{\text{جر ج مولية}}{\text{مول لیت}} \quad A_2 = 10 - 2, 19 = 7, 8 \text{ gr}$

$\frac{V \times 10^{-3} \times 100}{1} = 7, 8 \%$



$\frac{\frac{9 \text{ gr}}{\text{مول لیت}} \times V}{29, \text{r} \times 2} = \frac{V \times 10^{-3}}{1 \times 2} \rightarrow V \times 10^{-3} \times 29, \text{r} \times 2 = 2 \times 9 \text{ gr}$

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$$\sum KNOR \rightarrow \sum K_r O + \sum N_r + \sum O_r$$

$$\downarrow$$

$$K_r g_r$$

$$\downarrow$$

$$\text{Lit}$$

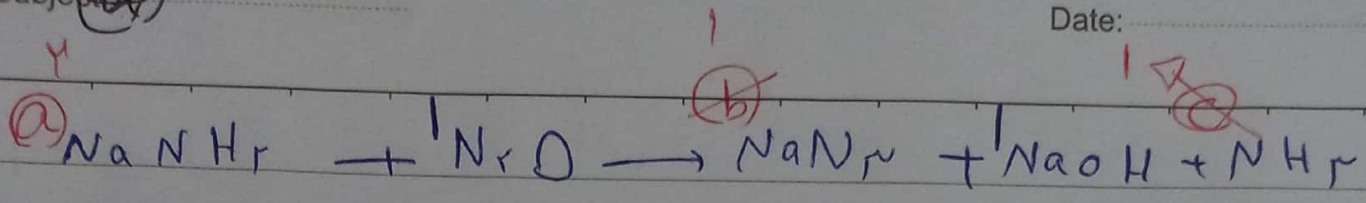
$$KNOR \sim O_r$$

$$\frac{P_r \cdot r_{gr}}{r \times 1.01} = \frac{V}{\delta \times \frac{1 \dots \times r_{gr}}{r \times 1.01}} \rightarrow V =$$

$$\frac{P_r V_r}{T_r} = \frac{P_r V_r}{T_r} \rightarrow \frac{1 \times r_{gr}}{r \times 1.01} = \frac{V_r}{\delta \times 1.01}$$

$$V = \frac{P_r \cdot r_{gr} \times \delta \times 1.01}{r \times 1.01} = \frac{P_r \cdot r_{gr} \times \delta}{r}$$

$$\approx 1.01 \cdot r_{gr}$$



$$\text{Na} \rightarrow a = b + 1 \rightarrow \boxed{b = a - 1}$$

$$\text{N} \rightarrow a + r = r b + c \rightarrow a + r = r a - r + c$$

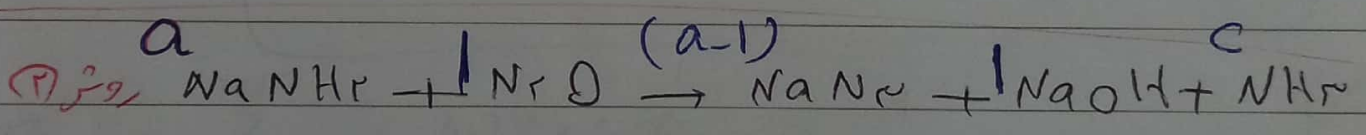
$$\boxed{0 = r a + c}$$

$$\text{H} \rightarrow \boxed{r a = 1 + r c} \quad | \quad r a = r a - r c$$

$$\boxed{r c = 1 \rightarrow c = 1}$$

$$\boxed{a = 2}$$

$$b = 1$$

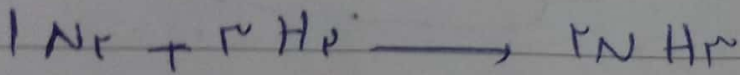
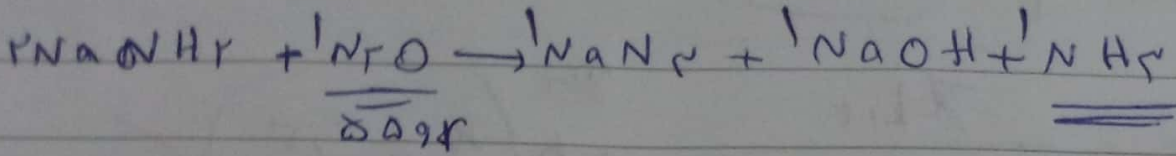


$$a = a - 1 + 1$$

$$\text{N} \rightarrow a + r = r a - r + c$$

$$\text{H} \rightarrow r a = 1 + r c \quad \rightarrow \quad a = r$$

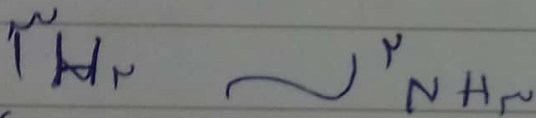
$$c = 1$$



ع. 1.3



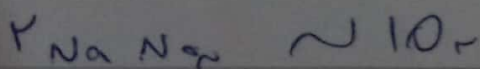
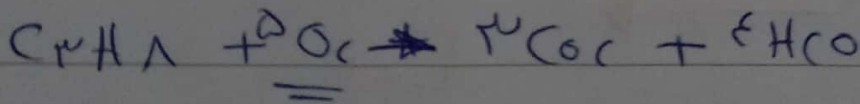
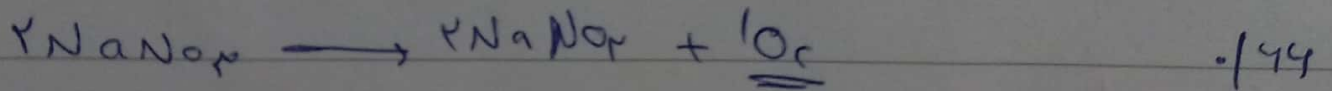
$$\frac{90 \times 22}{100 \times 1 \times 22} = \frac{x}{1} \rightarrow \boxed{\frac{22}{2} \times \frac{9}{10} = \text{mol}} \Rightarrow$$



ع. 1.3

~~$$\frac{90 \times 22}{100 \times 1 \times 22} = \frac{x}{1}$$~~

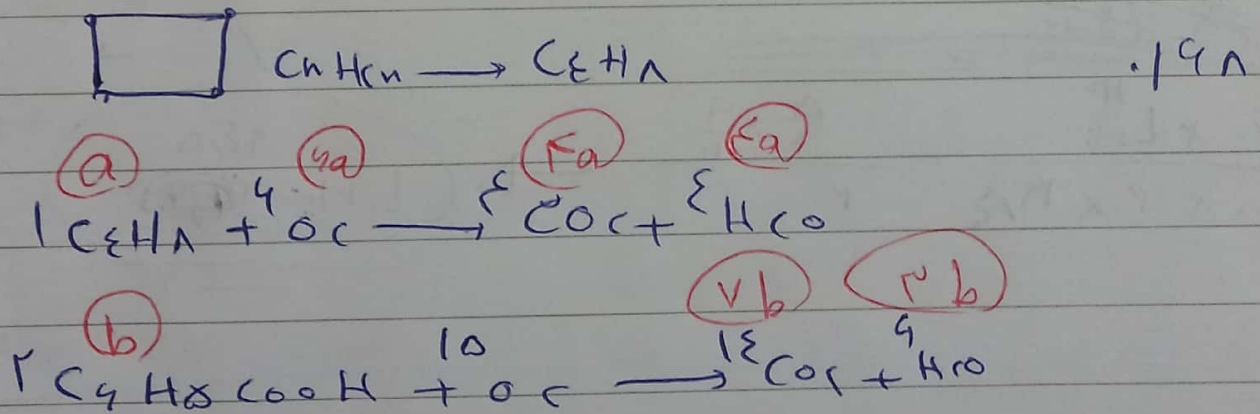
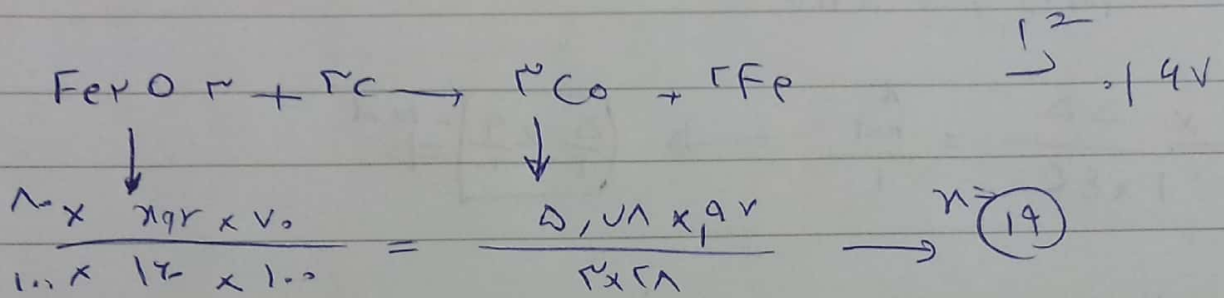
$$\frac{90 \times 22}{100 \times 1 \times 22} = \frac{x}{1} \rightarrow \boxed{\frac{22}{2} \times \frac{9}{10} = \text{mol}} \Rightarrow$$



$$\frac{100 \times 22 \times 22}{100 \times 1 \times 22} = \frac{\text{mol}}{1} \rightarrow \frac{22 \times 22 \times 22}{100 \times 1 \times 22} = \frac{100}{100} \text{ mol}$$

$$\Delta O_r \sim \sqrt{Lit \cdot i_b^v}$$

$$\frac{11 \frac{Lit}{No}}{\cancel{0} \times \cancel{0}} = \frac{\cancel{Lit}}{v \times \cancel{0} \cdot \Delta} \rightarrow \boxed{x = \epsilon, \tau} \checkmark$$



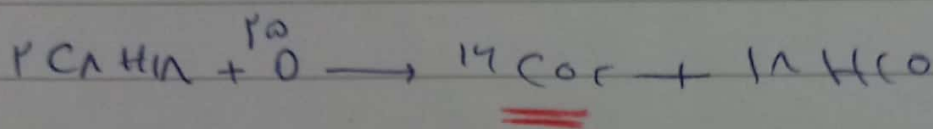
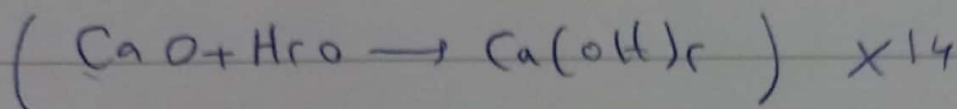
$$\begin{aligned} \tau a + 4b &= \tau \tau, \epsilon & a &= \tau / 1 \\ \tau a + 10b &= 1 \epsilon, \epsilon & b &= \tau \end{aligned}$$

$$\frac{\tau}{\epsilon / 1} \times 100 = \epsilon \cdot 11 /$$

$$\tau \times \tau / 1 \times 1A = 101 \tau$$

سوئال 149

جواب



$$r CaH_2 \sim 14 CaO$$

$$\frac{14 \times 17 \times 1}{r \times 112} = \frac{x}{14 \times 28} \rightarrow x = 32.5$$

$$CaH_2 \sim Ca(OH)_2$$

$$\frac{14 \times 17}{r \times 112} = \frac{mol}{14} \rightarrow mol = 4.25\%$$

$$CaH_2 \sim CO_2$$

$$\frac{14 \times 17}{r \times 112} = \frac{x}{14} \rightarrow 4.25 \text{ mol } CO_2$$

$$CO_2 \rightarrow CaO$$

$$\frac{4.25}{1} = \frac{x}{1} \rightarrow 4.25$$

$$\frac{CaO}{4.25} \sim \frac{Ca(OH)_2}{x}$$