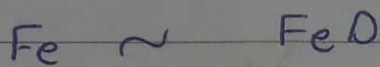
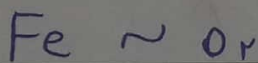


مرحلة 1

$$\frac{\text{كتلة}}{\text{عدد ذرات}} = \frac{9r}{\text{عدد ذرات}} = \frac{9r}{\text{عدد ذرات}}$$

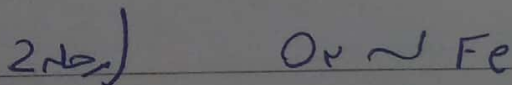


$$\frac{\frac{A}{T} \times 28}{8 \times 56} = \frac{2}{3 \times 32} \rightarrow \underline{\underline{x = 1,9 \text{ gr}_O}}$$



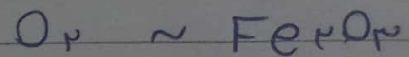
$$\frac{\frac{A}{T} \times 28}{8 \times 56} = \frac{y}{2 \times 32} \rightarrow y = 9,4 \text{ gr}_O$$

المسألة 2)  $\rightarrow 9,4 + 1,9 = 11,3$



$$\frac{1,9}{32} = \frac{y}{56}$$

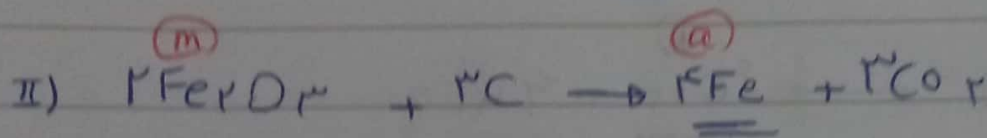
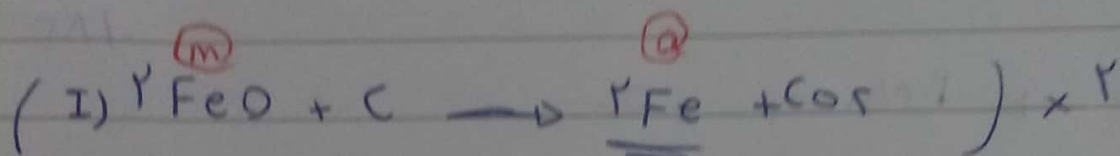
$$\rightarrow y = 1 \text{ mol}$$



$$\frac{9,4}{32} = \frac{x \text{ mol}}{2}$$

$$\rightarrow x = 1 \text{ mol}$$

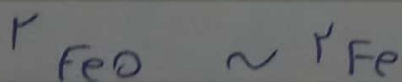
$$\frac{\text{mol } Fe_2O_3}{\text{mol } Fe} = \frac{1}{1} = \textcircled{1}$$



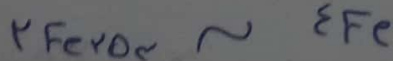
$\frac{P_1}{1} \times 96 = \frac{P_2}{1} \times 56$

$P_1 \frac{P_1 \times m}{E \times VC} = \frac{P_2 \times m}{r \times 14}$

$\rightarrow \frac{P_1}{P_2} = \frac{9}{14}$



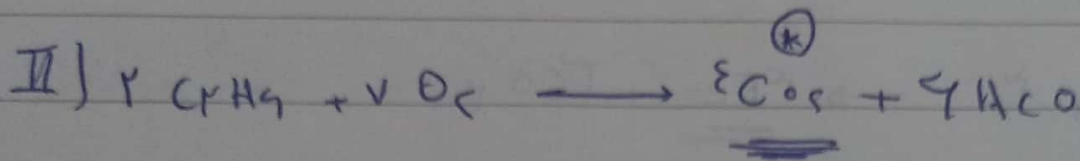
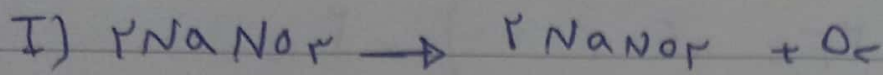
(2)  $\frac{P_1 \times m}{1 \times r \times VC} = \frac{9}{r}$



$\frac{P_2 \times m}{1 \times r \times 14} = \frac{9}{E}$

$\rightarrow \frac{P_1}{P_2} = \frac{9}{14}$

-112

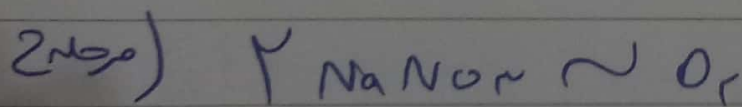


$$\varepsilon \text{CO}_2 \sim \nu \text{O}_2$$

$$\frac{\nu}{\varepsilon \times 32} = \frac{x_{\text{mol}}}{\nu} \rightarrow x = \frac{\nu}{\varepsilon} \text{ mol O}_2$$

$$\frac{\frac{\nu}{100} \text{ mol O}_2}{100} = \text{mol O}_2$$

$$\frac{\nu}{100} \times y = \frac{\nu}{\varepsilon} \rightarrow \frac{\nu}{\varepsilon} \text{ mol O}_2$$



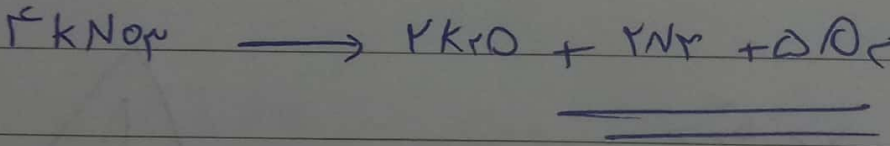
$$\frac{M \times x_{\text{gr}}}{100 \times \nu \times \Lambda} = \frac{\nu}{\varepsilon} \text{ mol} \rightarrow x = \frac{M \cdot \nu}{\varepsilon} \text{ gr}$$

$KNO_3 \sim 2 \text{ یون}$

$$\frac{12, 12}{1 \times 1.1} = \frac{\text{یون}}{2 \times NA} \rightarrow \boxed{1, 2 \times NA}$$

$Al_2O_3 \sim 3 O^{2-}$

$$\frac{? \text{ gr}}{1.2} = \frac{1, 2 \times NA}{2 \times NA} \Rightarrow \boxed{1, 19} \checkmark \quad \xi^2$$



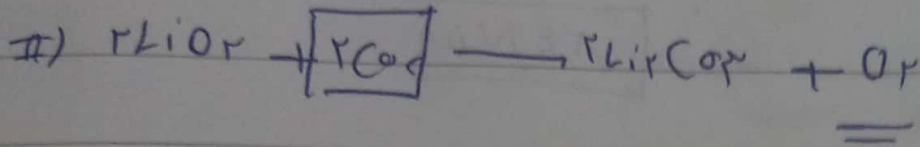
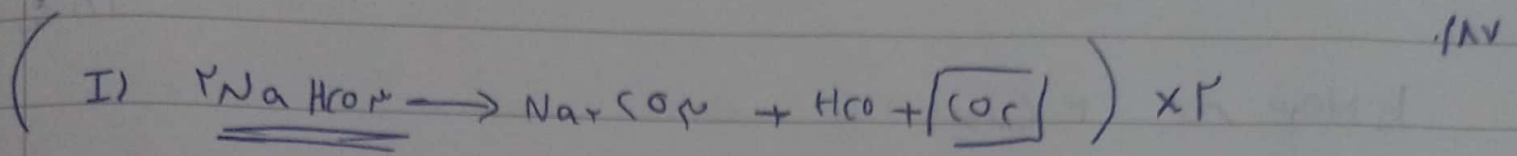
جرم خالص = جرم خالص - جرم خالص

$2 \times 112 = 20 - 6 \rightarrow 112 \times 2 = 112 \times 2 \text{ gr}$

$$\frac{\text{gr}}{2 \times 112} = \frac{112 \times 2}{2 \times 112} \rightarrow \text{gr}$$

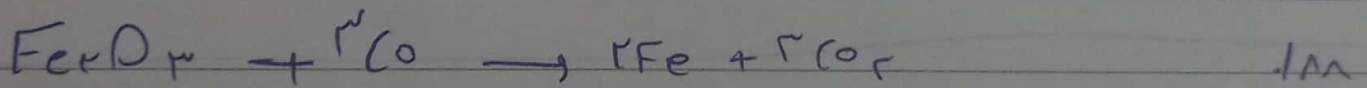
$2 N_2, O_2 \sim 2 KNO_3$

$$\frac{112 \times 2}{2 \times 112 \times 2} = \frac{2}{2 \times 112} \rightarrow 2 = \text{gr}$$



$\leftarrow Na_2HCO_3 \sim O_2$

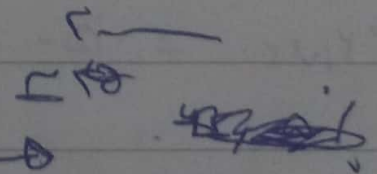
$$\frac{1.5 \times 10^3}{8 \times 10^3} = \frac{V_{LiOH}}{1 \times 2 \times 10^3} \rightarrow 11.25$$

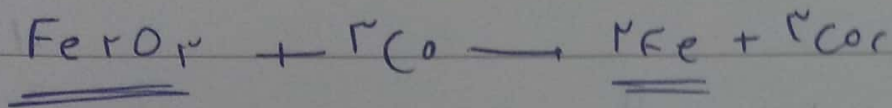


المعادلة 1292  $\rightarrow$  677:677

$1Fe_2O_3 \sim 3CO_2$

$$1 \times \frac{160}{100 \times 100} = \frac{12}{33 \times 100}$$



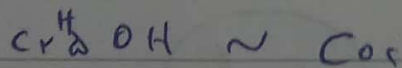
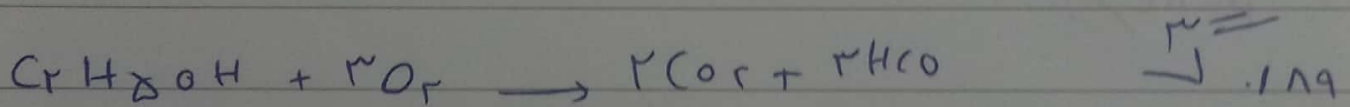


$$f_{Fe_2O_3} \rightarrow \left| -1 \times 14 + \nu \times 24 \right| \Rightarrow \textcircled{FA}$$

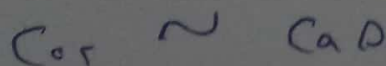
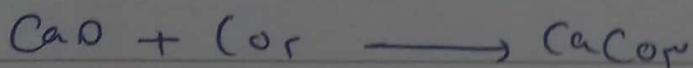
$$\frac{Fe_2O_3}{\nu_{Fe_2O_3}} = \frac{f_{Fe_2O_3}}{|\textcircled{FA}|}$$

$$\frac{P \times 48}{1 \times 12} = \frac{14}{|\textcircled{FA}|}$$

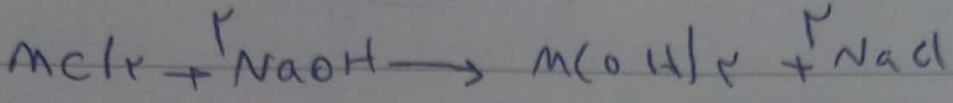
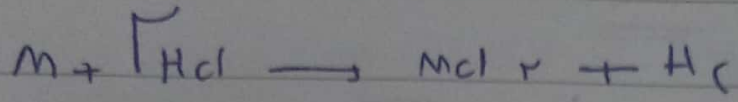
$$\rightarrow P = 95.12\%$$



$$\frac{92.92}{1 \times 44} = \frac{\text{mol}_{CO_2}}{\nu} \rightarrow \textcircled{f_{mol_{CO_2}}}$$



$$\frac{f_{mol}}{1} = \frac{92.92}{44} \rightarrow \textcircled{\lambda = 2.11}$$



$M \sim H_r$

$\frac{gr}{\text{حجم محلول}} = \frac{\text{النسبة}}{\text{حجم محلول}}$

$$\frac{P \times 100}{100 \times 1 \times 24} = \frac{rC, 9}{1 \times 100 \times 2}$$

$\rightarrow P = 12\%$

$H_r \sim rNaOH$

نسبة مolar

$$\frac{rC, 9}{1 \times 100 \times 2} = \frac{m_{mol}}{r}$$

$\rightarrow$   $r_{mol}$

$\xi^2$